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potassium alum. The synthesis requires the careful transfer solutions and some evaporation and cooling techniques.

Major works Estimated Time En route (min)

Reaction of aluminum foil with 4 M KOH and Filtration 10-15

Addition of 6 M H<sub>2</sub>SO<sub>4</sub> 4-5 Heating for

Dissolution of Al(OH)<sub>3</sub>

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with stirring 2030  
Crystallization 40  
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Suction Filtration and  
...

**Objectives**  
**Introduction Result**  
**Videos Report Sheet**

...  
Synthesis and Analysis  
of Potassium  
Aluminium Sulphate  
(Alum) from Waste  
Aluminium Can  
International Journal of  
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## Potassium Alum

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(IJARCS) Page 3 2

$Al(OH)_3(s) + 6 H^+(aq) \rightarrow 2 Al^{3+}(aq) + 6 H_2O(l)$  (6) The solution at this point contains  $Al^{3+}$  ions,  $K^+$  ions (from potassium hydroxide), and  $SO_4^{2-}$  ions

## **Synthesis and Analysis of Potassium Aluminium Sulphate**

...

In this lab activity we will synthesize

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potassium alum from aluminum metal. The synthesis is a multi-step process; first, aluminum metal dissolves by reaction with potassium hydroxide to form the aluminate complex ion:

$$2\text{Al}(s) + 2\text{K}^+(aq) + 2\text{OH}^-(aq) \rightarrow 2\text{AlO}_2^-(aq) + \text{H}_2(g)$$

## **Inorganic Synthesis: Preparation of Potassium Alum**

Lab report on synthesis of Alum using Aluminum, 1. can into

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## Potassium Alum

### Synthesis Lab

the chemical compound potassium aluminum sulfate,  $KAl(SO_4)_2 \cdot 12 H_2O$ , commonly referred to as alum. Data and Calculations: In this lab, a beverage can is used to get aluminum sheet and then, sheet

### **Lab report on synthesis of Alum using Aluminum.**

The name alum commonly refers to potassium aluminum



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## Potassium Alum

### Synthesis Lab

sulfate dodecahydrate,  $KAl(SO_4)_2 \cdot 12H_2O$ , which is a crystalline white solid. Among the many uses for alum are water purification, leather tanning, mordant dyeing, and as a component in baking powder. A synthesis may involve one chemical reaction or a series of chemical reactions.

## **Experiment 7:**

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**Synthesis of Alum**

Synthesis of Potassium

Aluminum Sulfate

Dodecahydrate Obtain

a piece of aluminum

foil weighing about 0.5

g and weigh it

precisely (to the

nearest 0.001g). Cut

the weighed foil into

many small pieces. The

smaller the pieces the

faster the reaction will

go because of the

increased surface area

exposed to the KOH

solution.

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## **Preparation and Analysis of Alum | Chem Lab**

Balance the overall equation for the synthesis of alum,  $KAl(SO_4)_2 \cdot 12 H_2O$ , from aluminum, potassium hydroxide, sulfuric acid and water. HINT: Add together the four equations given in the introduction, canceling out any molecules or formula units that appear on

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both sides.  $\text{Al} + \text{KOH} + \text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{KAl}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O} + \text{H}_2 \uparrow$

## **Synthesis of Alum from Aluminum**

The Synthesis of Alum  
Lab Michaela Tonsager  
and Kaili Johnson

Conclusion We  
determined that our  
sample was in fact  
alum. Our melting  
point of 99.4 degrees C  
was similar to the  
published melting point

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of 92.5 degrees C. Our percent sulfate was 42.44%, which is close to the

## **The Synthesis of Alum Lab by Michaela Tonsager**

Synthesis of a Hydrate:  
Alum is a general name for the type of compound in which many combinations of an alkali metal and ammonium or a trivalent metal such as aluminum, iron, or

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## Potassium Alum

### Synthesis Lab

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chromium. A synthesis reaction is a reaction in which two or more chemicals are combined to create a new compound or compounds.

### **The Formula, Synthesis, and Analysis of Alum - Odinity**

2 b) Preparation of Potassium hexathiocyanatochromate(III),  $K_3[Cr(NCS)_6]$  Make an aqueous solution of

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## Potassium Alum

### Synthesis Lab

potassium thiocyanate, KSCN (2.5 g), chrome alum (  $\{KCr(SO_4)_2 \cdot 12H_2O\}$  3.0 g) using distilled water (10 cm<sup>3</sup>). Pour the solution into an evaporating dish and place on a steam bath. Evaporate to dryness, to obtain a mass of red crystals.

## **CHEM2111**

### **Laboratory**

### **Experiments**

However, since we

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## Potassium Alum

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proved the identities of the proposed end product as those of Alum, we have successfully synthesizes Potassium Aluminum Sulfate from Aluminum scrap by the dissolution of Al (s) in a re-dox reaction.

Synthesis of Potassium Aluminum Sulfate (Alum) by Chemical Recycling of Scrap Aluminum The Importance of Aluminum



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## **Synthesis of Potassium Aluminum Sulfate (Alum) by by**

...

The overall reaction to form alum is:  $\text{Al}^{3+}(\text{aq}) + \text{K}^{1+}(\text{aq}) + 2 \text{SO}_4^{2-}(\text{aq}) + 12 \text{H}_2\text{O} \rightarrow \text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}(\text{s})$  Crystallization.

Using tongs, remove the filtrate beaker to a wire gauze on the lab bench. Allow it to begin to cool while you prepare an ice bath, an

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ice-water mixture in a  
400-mL beaker.  
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## **Synthesis of Alum - Web.LeMoyne.Edu**

In this video we will be synthesising potassium alum ( potash alum , syn- phitkari (hindi)) Potassium aluminium sulfate It is a double salt composed of Potassium sulfate Aluminium sulfate with 24 ...

**Synthesis of**  
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**Potassium alum  
(Potassium  
Aluminium  
Sulfate) NCERT**

Analysis of alum Part 2:  
Determination of the  
Water of Hydration in  
Alum Crystals Analysis  
of alum Trial 1 Trial 2  
Trial 3 Mass of crucible  
and cover (g) 45.9447  
41.5092 43.2373 Mass  
of crucible, cover, and  
alum crystals (g)  
47.9482 43.5127  
45.2412 Mass of alum  
crystals (g) 2.0035

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2.0035 2.0039  
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**Analysis of Alum,  
 $\text{AlK}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O}$**

For the synthesis of potassium alum, advantage is taken by the fact that aluminum hydroxide is amphoteric. An aluminum can is cut into small pieces. A 62g sample of the aluminum chips is used to prepare the potassium alum according to the

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procedure described in  
the experiment.

## **Experiment 17** **Prelab Assignment:** **Synthesis of an** **Alum ...**

Synthesis of Alum Alum is a solid ionic compound with many uses. It is used as astringent to prevent bleeding from small cuts, as an ingredient in deodorants, as an ingredient in baking powders, and as a

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## Potassium Alum

### Synthesis Lab

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preservative used in pickling. The formula of alum is  $KAl(SO_4)_2 \cdot 12H_2O$  and the standard name is potassium aluminum sulfate. It is an

## **CHEM 231**

### **Experiment 5**

#### **Synthesis of Alum**

The major sources of error in the synthesis of alum from aluminum foil include loss of product through various means, human

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and systematic errors,  
contamin

## **What Are Some Sources of Error in Synthesis of Alum From ...**

More time spent in lab,  
more careful  
measurements, and  
the elimination of  
impurities would have  
dictated a more  
precise yield. A cloudy  
participate was made  
when barium chloride  
was added to alum and

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the test indicates the presence of SASS- in the form of Bases.

**Synthesis Of Alum  
Essay Example -  
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Experiment 7:  
Synthesis of Alum Lab  
#4 Synthesis of Alum  
Purpose-Synthesize a  
type of alum called  
potassium aluminum  
sulfate dodecahydrate-  
Observe and record the  
process of  
synthesizing, and



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calculate the percent yield of the synthesis  
Procedure 1. Wear goggles 2. Obtain a small piece of aluminum foil and measure its mass using the analytical ...

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